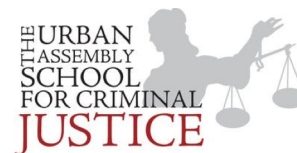


Name: _____ Date: _____

Chemistry ~ Ms. Hart

Class:

Anions or Cations



Use all of these words to make your concept map

- | | | |
|---------------------------|---------------------------|-----------------------|
| 1. Bright Line Spectra | 9. Group | 17. Noble Gases |
| 2. Bohr Model | 10. Period | 18. Cation |
| 3. Energy Level | 11. Metals | 19. Anion |
| 4. Electron Configuration | 12. Nonmetals | 20. Atomic Radius |
| 5. Ground State | 13. Metalloids | 21. Ionic Radius |
| 6. Excited State | 14. Alkali Metals | 22. Ionization Energy |
| 7. Valence Electrons | 15. Alkaline Earth Metals | 23. Electronegativity |
| 8. Lewis Dot Diagrams | 16. Halogens | |

Review Questions (complete these on a sheet of loose leaf)

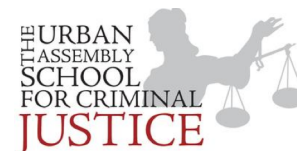
1. What is an excited state configuration for an atom of carbon-14?
2. What causes an atom to emit light?
3. What does it mean to have a completely full valence shell?
4. What is the ground state of an atom?
5. True/False. Metals have more valence electrons than nonmetals.
6. What is the trend in number of valence electrons as you move down a group?

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7. What is the Lewis Dot diagram for Nitrogen?
8. What elements are the metalloids?
9. List some characteristics of metals. How are these different from non-metals?
10. What determines the chemical properties of an element?
11. Which element on the periodic table has the lowest electronegativity value?
12. _____ is the tendency to gain electrons.
13. _____ is the _____ required to remove a valence electron.
14. Examine the relationship between atomic radius and ionic radius for metals. Compare this to the relationship between nonmetals.

Good luck studying for tomorrow's exam! Review all past concepts as there may be previous material on the test! Get a good nights sleep, use the website and contact me with questions!

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