Name:

MC Practice #2

1. What is the total number of electrons shared in a double covalent bond?	5. Which unit can be used to express the concentration of a solution?
(A) 1 (B) 2 (C) 3 (D) 4 2. Given the balanced equation representing a reaction:	(A) L/s(B) J/g(C) ppm(D) kPa6. Which formula represents a mixture?
$Br_2 + energy \rightarrow Br + Br$ Which statement describes the energy change and bonds in this	 (A) C₆H₁₂O₆(ℓ) (B) C₆H₁₂O₆(s) (C) LiCl(aq) (D) LiCl(s) 7. Which sample has particles with the <i>lowest</i> average kinetic energy?
 reaction? (A) Energy is released as bonds are broken. (B) Energy is released as bonds are formed. (C) Energy is absorbed as bonds are broken. (D) Energy is absorbed as bonds are formed. 3. Which substance can <i>not</i> be broken down by a chemical change? (A) methane (B) propanal (C) tungsten (D) water 4. Object A at 40°C and object B at 80°C are placed in contact with each other. Which statement describes the heat flow between the objects? (A) Heat flows from object A to object B. (B) Heat flows from object B to object A. (C) Heat flows in both directions between the objects. (D) No heat flow occurs between the objects. 	 (A) 1.0 g of I₂ at 50.°C (B) 2.0 g of I₂ at 30.°C (C) 7.0 g of I₂ at 40.°C (D) 9.0 g of I₂ at 20.°C 8. Which gas sample at STP has the same total number of molecules as 2.0 liters of CO₂(g) at STP? (A) 5.0 L of CO₂(g) (B) 2.0 L of Cl₂(g) (C) 3.0 L of H₂S(g) (D) 6.0 L of He(g) 9. Petroleum can be separated by distillation because the hydrocarbons in petroleum are (A) elements with identical boiling points (B) elements with different boiling points (C) compounds with different boiling points (D) compound is insoluble in water? (A) KOH (B) NH4Cl (C) Na3PO4 (D) PbSO4