

Unit 9

NAME

Class Work

4/17/14

## 9.6 Titrations

SPARK

Complete the SPARK on your worksheet!

## Objective

SWBAT calculate the concentration of acid/ base needed to neutralize a rxn.

# Answers to REVIEW sheet

1. 3

2. 3

3. 2

4. 2

5. 1

6. 4

7. 4

8. 4

9. 2

10. 1

11. 1

12. 2

13. 1

14. 1

15. 1

16. 4

17. 1

18. 1

19. yellow

20. —stomach—the  
organ with a pH of  
2

21. yellow

Objective: SWBAT calculate the concentration of acid/ base needed to neutralize a rxn.

# Titration

## Things to remember:

- A solution is neutral when the pH is \_\_\_\_\_.
- Molarity (M= “molar”) is a unit for \_\_\_\_\_.

# Think About This

- If a 1 liter solution of 1.5 M HCl needs to be neutralized, how many liters of a 1.5 M NaOH solution is needed?

# Titration:

- **TITRATION** is a process in which a specific volume of an acid or base with a known concentration is added to an acid or base of an unknown concentration until neutralization occurs.

# Demo

- Unknown concentration of acid
- Known concentration of base

# Let's Calculate it!

- Find the titration formula on your reference sheets!
- Switch to doc camera!

# HOMEWORK

Study for QUEST tomorrow!

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