

Name: _____ Date: _____

Chemistry ~ Ms. Hart

Class:

Anions or Cations



8.3 Classwork

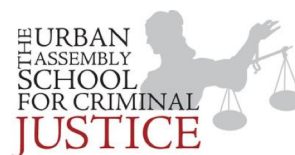
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| <p>1. Which ion, when combined with chloride ions, Cl^-, forms an insoluble substance in water?</p> <ul style="list-style-type: none">(1) Fe^{2+}(2) Mg^{2+}(3) Pb^{2+}(4) Zn^{2+} <p>2. According to Table F, which of these salts is <i>least</i> soluble in water?</p> <ul style="list-style-type: none">(1) LiCl(2) RbCl(3) FeCl_2(4) PbCl_2 <p>3. Which compound is insoluble in water?</p> <ul style="list-style-type: none">(1) BaSO_4(2) CaCrO_4(3) KClO_3(4) Na_2S | <p>4. According to Reference Table F, which of these compounds is the least soluble in water?</p> <ul style="list-style-type: none">(1) K_2CO_3(2) $\text{KC}_2\text{H}_3\text{O}_2$(3) $\text{Ca}_3(\text{PO}_4)_2$(4) $\text{Ca}(\text{NO}_3)_2$ <p>5. At standard pressure, a certain compound has a low boiling point and is insoluble in water. At STP, this compound most likely exists as</p> <ul style="list-style-type: none">(1) ionic crystals(2) metallic crystals(3) nonpolar molecules(4) polar molecules <p>6. Based on Reference Table F, which of these salts is the best electrolyte?</p> <ul style="list-style-type: none">(1) sodium nitrate(2) magnesium carbonate(3) silver chloride(4) barium sulfate |
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| <p>1. Which ion, when combined with chloride ions, Cl^-, forms an insoluble substance in water?</p> <ul style="list-style-type: none">(1) Fe^{2+}(2) Mg^{2+}(3) Pb^{2+}(4) Zn^{2+} <p>2. According to Table F, which of these salts is <i>least</i> soluble in water?</p> <ul style="list-style-type: none">(1) LiCl(2) RbCl(3) FeCl_2(4) PbCl_2 <p>3. Which compound is insoluble in water?</p> <ul style="list-style-type: none">(1) BaSO_4(2) CaCrO_4(3) KClO_3(4) Na_2S | <p>4. According to Reference Table F, which of these compounds is the least soluble in water?</p> <ul style="list-style-type: none">(1) K_2CO_3(2) $\text{KC}_2\text{H}_3\text{O}_2$(3) $\text{Ca}_3(\text{PO}_4)_2$(4) $\text{Ca}(\text{NO}_3)_2$ <p>5. At standard pressure, a certain compound has a low boiling point and is insoluble in water. At STP, this compound most likely exists as</p> <ul style="list-style-type: none">(1) ionic crystals(2) metallic crystals(3) nonpolar molecules(4) polar molecules <p>6. Based on Reference Table F, which of these salts is the best electrolyte?</p> <ul style="list-style-type: none">(1) sodium nitrate(2) magnesium carbonate(3) silver chloride(4) barium sulfate |
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7. Which barium salt is insoluble in water?
- (1) BaCO_3
 - (2) BaCl_2
 - (3) $\text{Ba}(\text{ClO}_4)_2$
 - (4) $\text{Ba}(\text{NO}_3)_2$
8. Based on Reference Table F, which of these saturated solutions has the lowest concentration of dissolved ions?
- (1) $\text{NaCl}(\text{aq})$
 - (2) $\text{MgCl}_2(\text{aq})$
 - (3) $\text{NiCl}_2(\text{aq})$
 - (4) $\text{AgCl}(\text{aq})$
9. At STP, which of these substances is most soluble in H_2O ?
- (1) CCl_4
 - (2) CO_2
 - (3) HCl
 - (4) N_2

10. Based on Reference Table F, describe the solubility of magnesium hydroxide in water.

11. Based on Reference Table F, describe the solubility of zinc sulfide in water.

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