Unit 8 NAME
Class Work 3/19/14

8.2 Factors Affecting Solubility

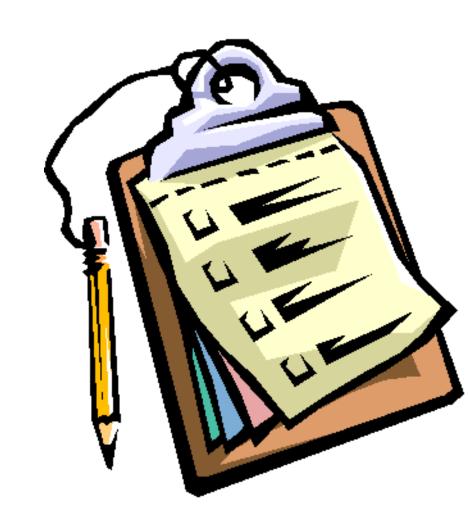
SPARK (take out 8.1)

- 1. List the factors that affect the rate of a reaction.
- 2. What are the two parts of a solution?

Objective

Agenda:

- SPARK/Objective
- Notes
- Practice
- Homework



Review-Solutions

Checklist:

- Homogeneous mixtures
- Clear and do not disperse light
- Can have color
- Do not settle on standing
- Pass through a filter

Notes

- Solubility**: how easily something dissolves
- Soluble** Something that has high solubility.
 E.g. Salt is soluble in water.
- Insoluble** Low solubility. E.g. plastic and water.

Factors Affecting Solubility

Factor	Effect	Exceptions
Temperature	Increase the temperature more soluble	GASES
Pressure	Does nothing	GASES Increase pressure of a gas that's in a liquid, you increase solubility
Polarity	Causes attraction btwn molecules of solute and solvent LIKE DISSOLVES LIKE	

Types of Solvents!

Polar Solvents

- Water, H₂O
- Ethanol, CH₃CH₂OH
- Acetone, C₃H₆O

Nonpolar solvents

- Hexane, C₆H₁₄
- Benzene, C₆H₆
- Toluene, C₇H₈

Draw an image to represent the second bullet point below:

- Greases, which are nonpolar, won't easily wash off our hands in water, which is polar.
- Soaps have one end that is polar, which allows the soap to dissolve in water. The other end of the soap is nonpolar, and grease will dissolve in it.

Fill in the chart below:

Solute Type	Nonpolar solvent	Polar Solvent
	Soluble	
Polar		
		soluble

Check your work!

Solute Type	Nonpolar solvent	Polar Solvent
Nonpolar	Soluble	insoluble
Polar	insoluble	soluble
Ionic	insoluble	soluble

Classwork

Complete your 8.2 classwork!

Lab #19

- How do we write a hypothesis?
- What is the independent variable?
- What is the dependent variable?
- What should our graph look like?
- What should be included in the conclusion?

HOMEWORK

Finish 8.2 Classwork

Complete the 8.2 HW sheet
Lab #19