

Unit 8

NAME

Class Work

3/17/14

## 8.1 Introduction to Solutions

SPARK (pass your HIV work in the HW bin)

1. What is the difference between a homogeneous solution and a heterogeneous solution?
2. Give one example of a homogeneous solution.
3. Give one example of a heterogeneous solution.

## Objective

SWBAT define solution.

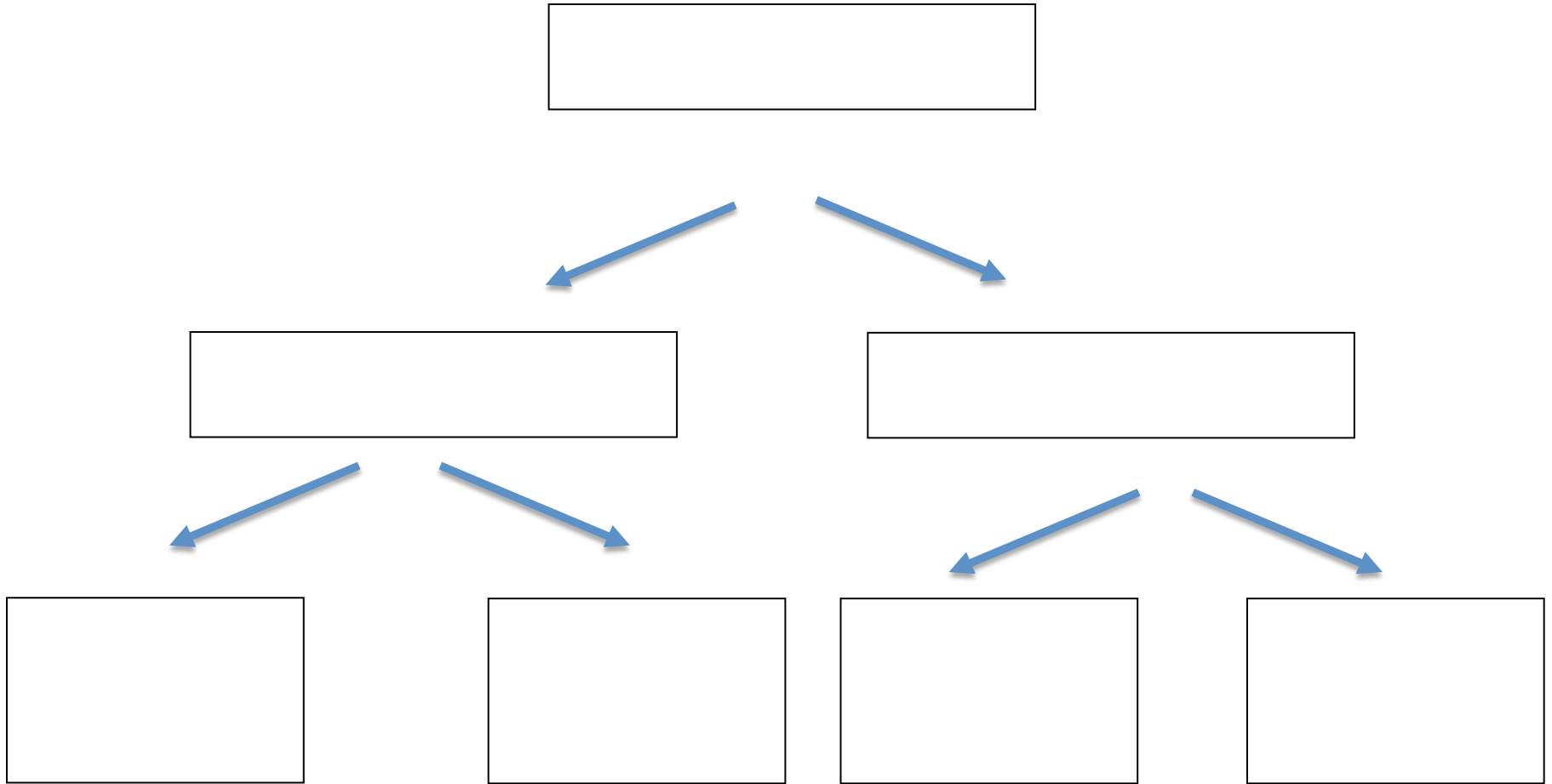
# Agenda:

- SPARK/Objective
- Notes
- Practice
- Homework



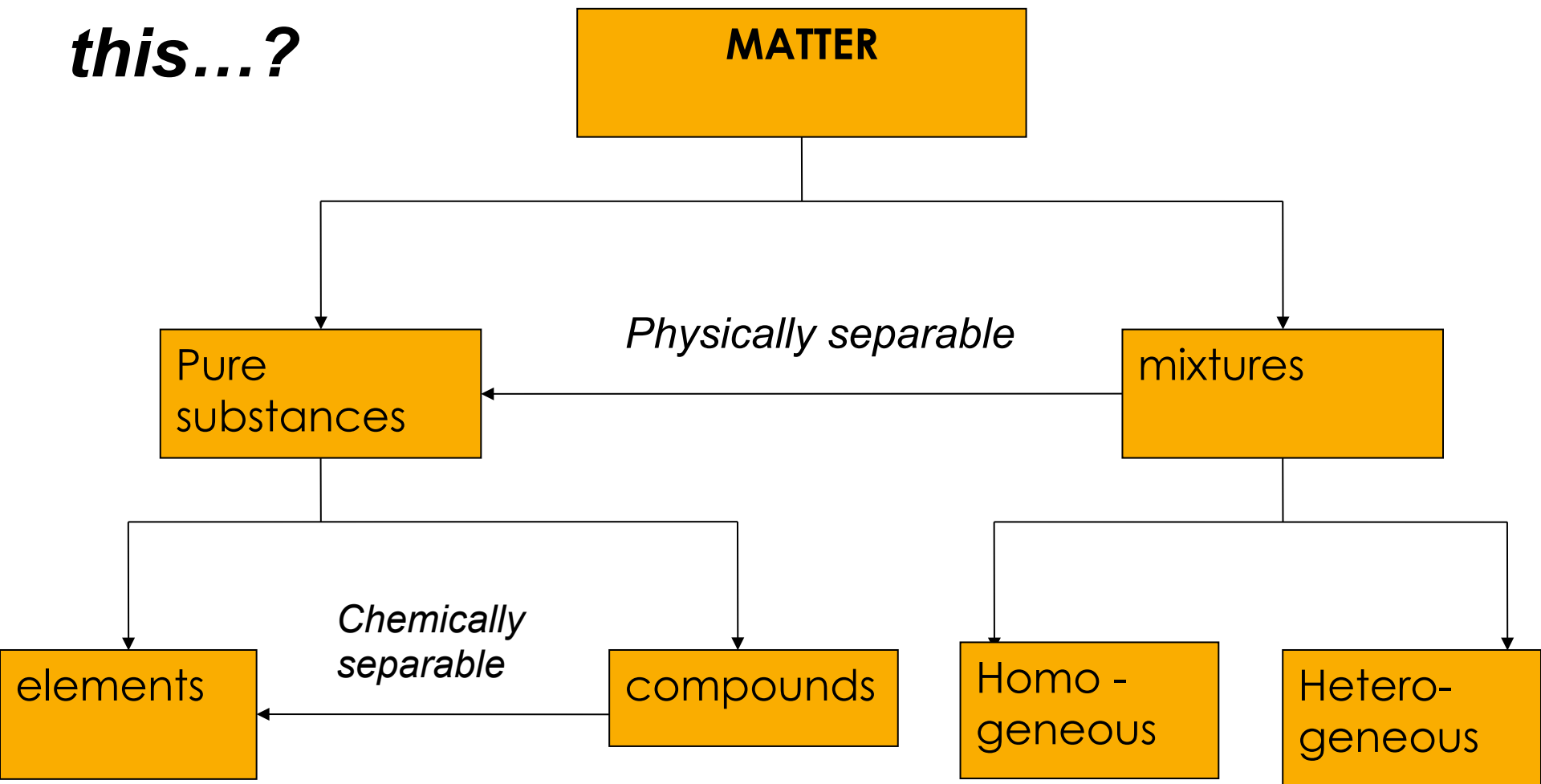
Objective: SWBAT define solution

# REVIEW- kicking it old school...



Objective: SWBAT define solution

***Remember  
this...?***



Objective: SWBAT define solution

# Quick Check!

- Which sample of matter is classified as a substance?
  1. Air
  2. Ammonia
  3. Milk
  4. Seawater

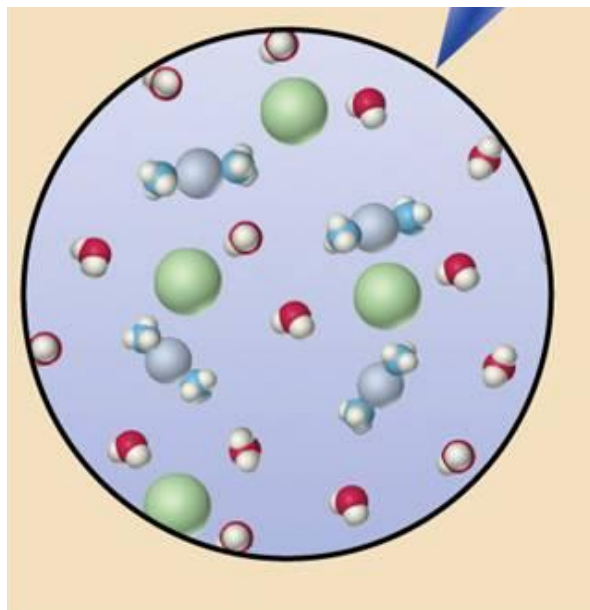
# Notes

- A solution\*\* is a homogeneous mixture of substances

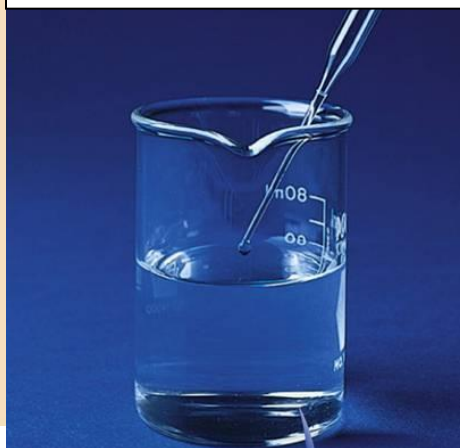


Objective: SWBAT define solution

# What does a solution look like?

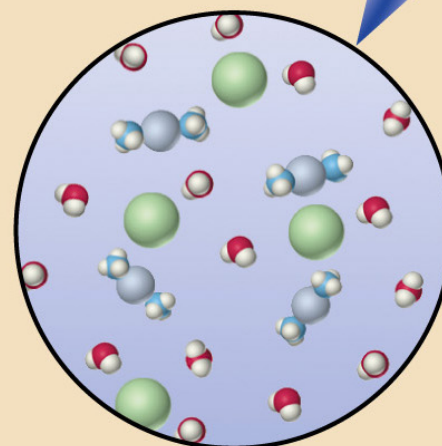
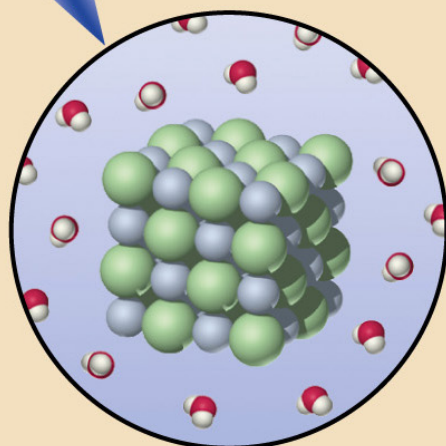
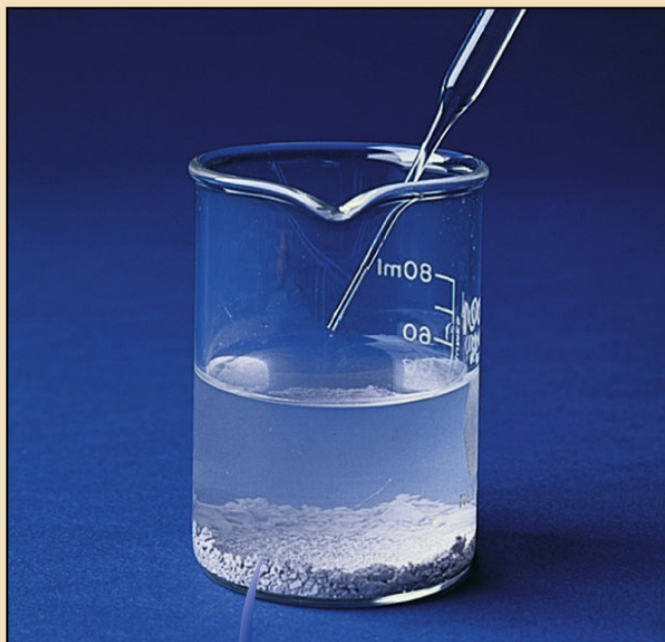


Liquid (NH<sub>3</sub>  
in H<sub>2</sub>O)



Solid  
(Metal alloy)  
Brass = Zn+Cu

Objective: SWBAT define solution

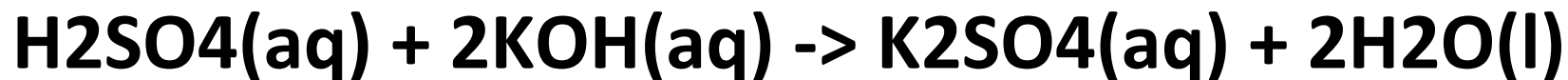


Objective: SWBAT define solution



# Notes

- A solution is also referred to as aqueous = (aq)



Objective: SWBAT define solution

# Notes

- There are 2 parts of solution:
  - 1) Solute – what is being dissolved
  - 2) Solvent – what is doing the dissolving

Objective: SWBAT define solution

# Practice

Solution	Solute	Solvent
NaCl (aq)		
C <sub>6</sub> H <sub>12</sub> O <sub>6</sub> (aq)		
AgNO <sub>3</sub> (aq)		

Objective: SWBAT define solution

# Classwork

- Complete your 8.1 classwork!

Objective: SWBAT define solution

# Complete your Character Survey!

- You have five minutes to reflect on this- place it in the classwork bin!

Objective: SWBAT define solution

# HOMEWORK

Complete the 8.1 WS!

Have parents complete character survey!

Objective: SWBAT define solution