Unit 7
Class Work

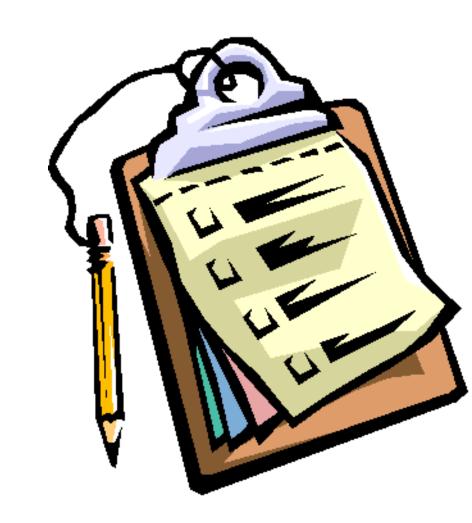
NAME 3/4/14

7.6 Entropy

SPARK
Complete your 7.5 Exit Ticket

# Agenda:

- SPARK/Objective
- Notes
- Practice
- Homework



### Review of 7.5 HW

- B
- 2. **D**
- 3. <u>C</u>
- 4. <u>C</u>
- 5. A
- 6. **B**
- 7. <u>A</u>
- 8. <u>D</u>

Acceptable responses include, but are not limited to: • A lower concentration of oxygen gas decreases the number of effective collisions between O<sub>2</sub> molecules and CH<sub>4</sub> molecules.

- 10. B
- 11. <u>C</u>
- 12. **B**
- 13. <u>C</u>
- 14. <u>B</u>

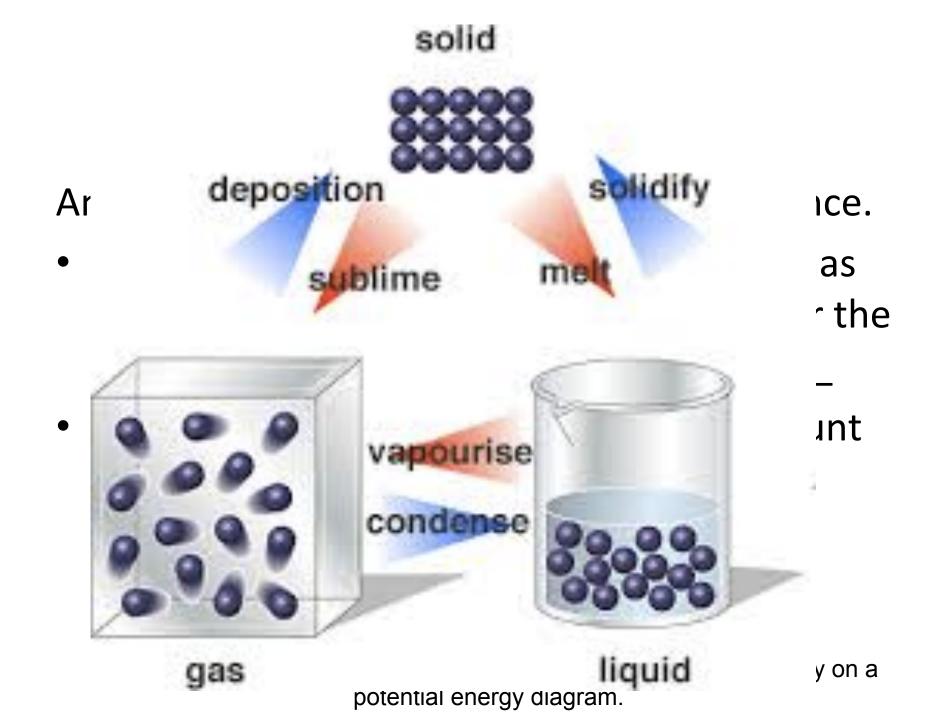
## **Notes - Entropy**

- Entropy is a measure of the RANDOMNESS or DISORDER of the substances in a reaction
- The greater the TEMPERATURE, the HIGHER THE ENTROPY.
- Draw the particle diagrams for the three phases of water. Use a solid dot to represent one molecule of water!

## Stop and Jot

Answer the following in a COMPLETE sentence.

- Which of the following phases of matter has the greatest entropy (hint: first remember the three states of matter)?
- Which phase of matter has the least amount of entropy?



## **Notes - Entropy**

 Systems tend to undergo changes to lower energy and higher entropy



### Practice

 Practice: State whether entropy increases or decreases.

<b>Process/ Reaction</b>	Increase or decrease in entropy?
Ice melting	
Organizing your chem notebook	
Sublimination	
Deposition	

# Show me your FINGERS!

#### 1. Given the balanced equation:

$$I_2(s) + energy \rightarrow I_2(g)$$

As a sample of  $I_2(s)$  sublimes to  $I_2(g)$ , the entropy of the sample:

- 1 increases because the particles are less randomly arranged
- (2) increases because the particles are more randomly arranged
- (3) decreases because the particles are less randomly arranged
- $\stackrel{\textstyle ullet}{\textstyle 4}$  decreases because the particles are more randomly arranged

### Classwork

Complete your 7.6 classwork!

## **Exit Ticket**

Complete

#### **HOMEWORK**

Complete the 7.6 Review Sheet!
Study for your Unit 7 TEST on Thursday!