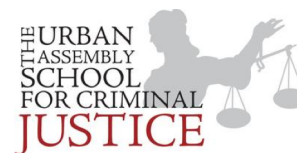


Name: _____ Date: _____

Chemistry ~ Ms. Hart

Class: _____ Anions or Cations



7.5 Enthalpy Exit Ticket

1. Use Table I to determine what the heat of formation is for the reaction of hydrogen and iodine to produce 2HI.

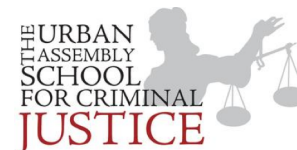
2. Given a ΔH of +66.4, is the reaction endothermic or exothermic? _____

3. Given the equation: $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \leftrightarrow 2\text{H}_2\text{O}(\text{l}) + 571.6 \text{ kJ}$
Determine the amount of heat released by the production of 4 mole of H_2O .

Name: _____ Date: _____

Chemistry ~ Ms. Hart

Class: _____ Anions or Cations



7.5 Enthalpy Exit Ticket*

1. Use Table I to determine what the heat of formation is for the reaction of nitrogen and oxygen to produce 2NO (note: no subscript 2).

2. Given a ΔH of -84.0, is the reaction endothermic or exothermic? _____

3. Given the equation: $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \leftrightarrow 2\text{H}_2\text{O}(\text{l}) + 571.6 \text{ kJ}$
Determine the amount of heat released by the combustion of 1 mole of H_2 .