Name: _____ **Date:** _____

Chemistry *~ Ms. Hart* **Class:** Anions or Cations

7.5 Enthalpy Exit Ticket

1. Use Table I to determine what the heat of formation is for the reaction of hydrogen and iodine to produce 2HI.

2. Given a ΔH of +66.4, is the reaction endothermic or exothermic? _____

3. Given the equation: 2H2(g) + O2(g) ↔ 2H2O(l) + 571.6 kJ
Determine the amount of heat released by the production of 4 mole of H_2O.

Name: _____ **Date:** _____

Chemistry *~ Ms. Hart* **Class:** Anions or Cations

7.5 Enthalpy Exit Ticket*

1. Use Table I to determine what the heat of formation is for the reaction of nitrogen and oxygen to produce 2NO (note: no subscript 2).

2. Given a ΔH of -84.0, is the reaction endothermic or exothermic? _____

3. Given the equation: 2H2(g) + O2(g) ↔ 2H2O(l) + 571.6 kJ
Determine the amount of heat released by the combustion of 1 mole of H_2.