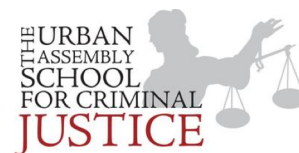


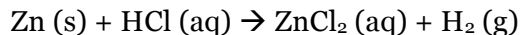
Name: _____ Date: _____

Chemistry ~ Ms. Hart Class: Anions or Cations



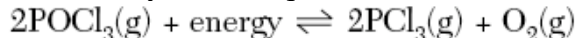
7.3 – Le Chatelier - Exit Ticket

1. Consider the following equation:



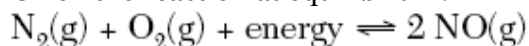
What will happen to the concentration of each product or reactant if ZnCl_2 is added to the system at equilibrium? **Place a + if the concentration will increase and a – if the concentration will decrease.**

2. Given the system at equilibrium:



Which changes occur when $\text{O}_2\text{(g)}$ is added to this system?

- (1) The equilibrium shifts to the right and the concentration of $\text{PCl}_3\text{(g)}$ increases.
 - (2) The equilibrium shifts to the right and the concentration of $\text{PCl}_3\text{(g)}$ decreases.
 - (3) The equilibrium shifts to the left and the concentration of $\text{PCl}_3\text{(g)}$ increases.
 - (4) The equilibrium shifts to the left and the concentration of $\text{PCl}_3\text{(g)}$ decreases.
3. Given the reaction at equilibrium:

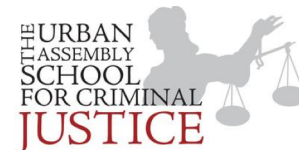


Which change will result in a *decrease* in the amount of NO (g) formed?

- (1) decreasing the pressure
- (2) decreasing the concentration of $\text{N}_2\text{(g)}$
- (3) increasing the concentration of $\text{O}_2\text{(g)}$
- (4) increasing the temperature

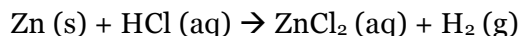
Name: _____ Date: _____

Chemistry ~ Ms. Hart Class: Anions or Cations



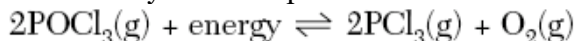
7.3 – Le Chatelier - Exit Ticket

1. Consider the following equation:



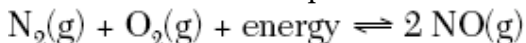
What will happen to the concentration of each product or reactant if ZnCl_2 is added to the system at equilibrium? **Place a + if the concentration will increase and a – if the concentration will decrease.**

2. Given the system at equilibrium:



Which changes occur when $\text{O}_2\text{(g)}$ is added to this system?

- (1) The equilibrium shifts to the right and the concentration of $\text{PCl}_3\text{(g)}$ increases.
 - (2) The equilibrium shifts to the right and the concentration of $\text{PCl}_3\text{(g)}$ decreases.
 - (3) The equilibrium shifts to the left and the concentration of $\text{PCl}_3\text{(g)}$ increases.
 - (4) The equilibrium shifts to the left and the concentration of $\text{PCl}_3\text{(g)}$ decreases.
3. Given the reaction at equilibrium:



Which change will result in a *decrease* in the amount of NO (g) formed?

- (1) decreasing the pressure
- (5) decreasing the concentration of $\text{N}_2\text{(g)}$
- (6) increasing the concentration of $\text{O}_2\text{(g)}$
- (7) increasing the temperature