Name:		Date:			HURBAN
Chemistry ~ Ms. Hart	Class:	Anions	or	Cations	FOR CRIM

FOR CRIMINAL **IUSTIC**

7.2 Chemical Equilibrium - Exit Ticket

- 1. What is required for a chemical reaction to occur?
 - (1) standard temperature and pressure
 - (2) a catalyst added to the reaction system
 - (3) effective collisions between reactant particles
 - (4) an equal number of moles of reactants and products
- 2. As the temperature of a chemical reaction in the gas phase is increased, the rate of the reaction increases because

- (1) fewer particle collisions occur
- (2) more effective particle collisions occur
- (3) the required activation energy increases
- (4) the concentration of the reactants increases
- 3. Which statement describes a chemical reaction at equilibrium?
 - (1) The products are completely consumed in the reaction.
 - (2) The reactants are completely consumed in the reaction.
 - (3) The concentrations of the products and reactants are equal.
 - (4) The concentrations of the products and reactants are constant.
- 4. Given the equation representing a closed system:
 - $N2O4(g) \leftrightarrow 2NO2(g)$

Which statement describes this system at equilibrium?

- (1) The volume of the NO2(g) is greater than the volume of the N2O4(g).
- (2) The volume of the NO₂(g) is less than the volume of the N₂O₄(g).
- (3) The rate of the forward reaction and the rate of the reverse reaction are equal.
- (4) The rate of the forward reaction and the rate of the reverse reaction are unequal.

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Chemistry ~ Ms. Hart	<u>Class:</u>	Anions or Cations	SCHOOL FOR CRIMINAL JUSTICE

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