**Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_**



***Chemistry*** *~ Ms. Hart* **Class:** Anions or Cations

**6.8 Chemical Reactions – Exit Ticket**

*Identify what type of reaction each of the following is – synthesis (S), decomposition (D), combustion (C), single-displacement (SD), or double-displacement (DD).*

* 1. K2CO3 + CuSO4 🡪 CuCO3 + K2SO4
	2. 2 NaHCO3 🡪 Na2CO3 + H2O + CO2
	3. H2O + SO3 🡪 H2SO4
	4. CaCO3 🡪 CaO + CO2
	5. Li2O + H2O 🡪 2 LiOH
1. Given the balanced equations representing two chemical reactions:

Cl2 + 2NaBr 🡪 2NaCl + Br2

 2 NaCl 🡪 2Na+ Cl2

Which types of chemical reactions are represented by these equations?

1. Single replacement and decomposition
2. Single replacement and double replacement
3. Synthesis and decomposition
4. Synthesis and double replacement
5. In which type of reaction do two or more substances combine to produce a single substance?
6. synthesis
7. decomposition
8. single replacement
9. double replacement

**Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_**



***Chemistry*** *~ Ms. Hart* **Class:** Anions or Cations

**6.8 Chemical Reactions – Exit Ticket**

*Identify what type of reaction each of the following is – synthesis (S), decomposition (D), combustion (C), single-displacement (SD), or double-displacement (DD).*

1. K2CO3 + CuSO4 🡪 CuCO3 + K2SO4
2. 2 NaHCO3 🡪 Na2CO3 + H2O + CO2
3. H2O + SO3 🡪 H2SO4
4. CaCO3 🡪 CaO + CO2
5. Li2O + H2O 🡪 2 LiOH
6. Given the balanced equations representing two chemical reactions:

Cl2 + 2NaBr 🡪 2NaCl + Br2

 2 NaCl 🡪 2Na+ Cl2

Which types of chemical reactions are represented by these equations?

1. Single replacement and decomposition
2. Single replacement and double replacement
3. Synthesis and decomposition
4. Synthesis and double replacement
5. In which type of reaction do two or more substances combine to produce a single substance?
6. synthesis
7. decomposition
8. single replacement
9. double replacement