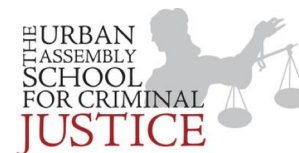


Name: _____ Date: _____

Chemistry ~ Ms. Hart

Class: _____ Anions or Cations



6.2 – Finding the Molecular Formula and Percent Composition

Exit Ticket - _____/4

Directions: Complete the following questions showing all work!

Vitamin C, also known as **ascorbic acid**, is water-soluble and cannot be produced by the human body. Each day, a person's diet should include a source of vitamin C, such as orange juice. Ascorbic acid has an empirical formula of $C_3H_4O_3$ and a gram-formula mass of **176 grams per mole**.

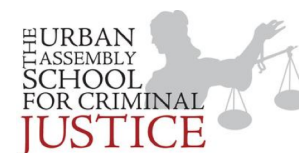
1. What is the **molecular formula** of ascorbic acid?

2. In the space *below*, **show a numerical setup and the calculated result** for **calculating the percent composition** by mass of **oxygen** in **ascorbic acid**. [1]

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