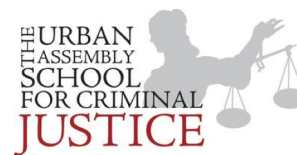


Name: \_\_\_\_\_ Date: \_\_\_\_\_

**Chemistry** ~ Ms. Hart

**Class:** \_\_\_\_\_ Anions or Cations



### 5.3 Naming Binary Ionic Compounds!

*Directions: If the name of the compound is given, write the correct chemical formula. If the formula is given, write the name (including the roman numeral!).*

1. KBr \_\_\_\_\_
2.  $V_2O_5$  \_\_\_\_\_
3. Cobalt (III) oxide \_\_\_\_\_
4. Barium phosphide \_\_\_\_\_
5. Cadmium nitride \_\_\_\_\_
6.  $Cu_3P$  \_\_\_\_\_
7.  $Ag_2S$  \_\_\_\_\_
8.  $Sn_3N_4$  \_\_\_\_\_
9. Radium iodide \_\_\_\_\_
10. Beryllium selenide \_\_\_\_\_
11.  $Fe_2S_3$  \_\_\_\_\_
12. SrO \_\_\_\_\_
13.  $CrCl_2$  \_\_\_\_\_
14. Mercury (II) fluoride \_\_\_\_\_
15. Lead (IV) bromide \_\_\_\_\_
16. CuSe \_\_\_\_\_
17. FeP \_\_\_\_\_
18. Lithium oxide \_\_\_\_\_
19. Cobalt (III) fluoride \_\_\_\_\_
20.  $CdI_2$  \_\_\_\_\_

21. Which element has an atom with the greatest tendency to attract electrons in a chemical bond?
- carbon
  - chlorine
  - silicon
  - sulfur
22. Which term indicates how strongly an atom attracts the electrons in a chemical bond?
- alkalinity
  - atomic mass
  - electronegativity
  - activation energy
23. Element X reacts with chlorine to form an ionic compound that has the formula  $\text{XCl}_2$ . To which group on the Periodic Table could element X belong?
- Group 1
  - Group 2
  - Group 13
  - Group 15
24. Which element forms an ionic compound when it reacts with lithium?
- K
  - Fe
  - Kr
  - Br
25. The bonds in BaO are best described as
- covalent, because valence electrons are shared
  - covalent, because valence electrons are transferred
  - ionic, because valence electrons are shared
  - ionic, because valence electrons are transferred
26. Which formula represents an ionic compound?
- $\text{H}_2$
  - $\text{CH}_4$
  - $\text{CH}_3\text{OH}$
  - $\text{NH}_4\text{Cl}$
27. Which substance contains bonds that involved the transfer of electrons from one atom to another?
- $\text{CO}_2$
  - $\text{NH}_3$
  - KBr
  - $\text{Cl}_2$
28. The correct chemical formula for iron(II) sulfide is
- FeS
  - $\text{Fe}_2\text{S}_3$
  - $\text{FeSO}_4$
  - $\text{Fe}_2(\text{SO}_4)_3$
29. The chemical formula for nickel (II) bromide is
- $\text{Ni}_2\text{Br}$
  - $\text{NiBr}_2$
  - $\text{N}_2\text{Br}$
  - $\text{NBr}_2$
30. What is the chemical formula for iron(III) oxide?
- FeO
  - $\text{Fe}_2\text{O}_3$
  - $\text{Fe}_3\text{O}$
  - $\text{Fe}_3\text{O}_2$

