Name:			Date:		- HURBAN
Chemi	stry ~ Ms. Hart	Class:	Anions of	r Cations	SCHOOL FOR CRIMINAL
	5.2 Ionic Bond	ing - Indepe	ndent Work		JUDITEL
Name the foll	owing <i>ionic</i> compou	nds:			
1) NaBr					
2) CaO_					
3) Li₂S_					
4) MgBr ₂	2				
Write the formulas for the following <i>ionic</i> compounds:					
5) potassium iodide					
6) magne	esium oxide				
7) alumi	num chloride				
		Pota	ssium (K) and F	luorine (F)	
8. Lewis Dot					
Structures with					
charges and arrows					
0 Chemical					
Equation					
		Rubi	dium (Rb) and H	Fluorine (F)	
10. Lewis					
Dot Structures					
with					
charges and					
11. Chemical					
Equation					

Use the Criss-Cross Method to write the chemical formula:

12. Calcium (Ca) and Oxygen (O)

13. Lithium (Li) and Iodine (I)

14. Beryllium (Be) and Sulfur (S)

15. Sodium (Na) and Bromine (Br)

Use the Box Method to write the chemical formula:

16. Potassium Fluoride

17. Magnesium Chloride

18. Barium Oxide

WHY DO ELEMENTS FALL IN LOVE?

Sodium (Na) metal is unstable with only 1 valence electron. It needs another element to help it become stable and happy. One day, it meets Chlorine (Cl) the nonmetal. Chlorine was also unstable and also needed another element to help it become stable and happy. It is love at first sight. When they meet, Sodium immediately gives its 1 valence electron to Chlorine so they can both be stable and happy. This gift of an electron to Chlorine makes them both charged and the force of attraction between their charges is too strong to resist so they fall in love and form an ionic bond, holding them together forever. The creation of their bond makes them overflow with energy and excitement. The closeness of their bond produces a love child, table salt (NaCl).

Why did the two elements need each other?

A Chemical Love Story:

What is produced by their bond?

What is this type of bond called?

What type of bond is made when a metal and a nonmetal react?