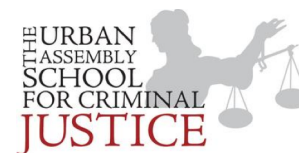


Name: _____ Date: _____

Chemistry ~ Ms. Hart

Class: _____ Anions or Cations



5.1 Introduction to Chemical Bonding

1. Which atom is most likely to bond with an atom?

- 1) An atom with 8 valence electrons. WHY?
- 2) An atom with 7 valence electrons.
- 3) An atom with 6 valence electrons.
- 4) An atom with 4 valence electrons.

2. Atom A has 7 electrons in its outer shell. Atom B has 1 electron in its outer shell. After bonding, both have 8 in their outer shell, but Atom A has a -1 charge and Atom B has a +1 charge. What kind of bond is this?

- 1) Ionic WHY?
- 2) Covalent
- 3) Metallic
- 4) Electronic

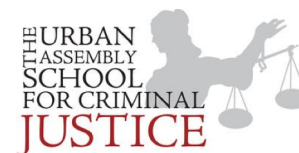
3. An atom has 15 total electrons, how many does it have in its outer shell?

- 1) 3 WHY?
- 2) 4
- 3) 5
- 4) 6

Name: _____ Date: _____

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- 2) 4
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4. Which atom is most likely to bond with another atom?
- 1) An atom with 8 valence electrons. WHY?
 - 2) An atom with 10 total electrons
 - 3) An atom with 2 total electrons
 - 4) An atom with 9 total electrons
5. Cations have positive charges. Anions have negative charges. What force draws the two together?
- 1) Magnetic Attraction WHY?
 - 2) Ionic Attraction
 - 3) Electromagnetic attraction
 - 4) Electrostatic attraction

Exit Ticket:

Answer the following in complete sentences: are most elements more stable as compounds or as elements? Explain.

*Use an example such as LiF to explain.

4. Which atom is most likely to bond with another atom?
- 5) An atom with 8 valence electrons. WHY?
 - 6) An atom with 10 total electrons
 - 7) An atom with 2 total electrons
 - 8) An atom with 9 total electrons
5. Cations have positive charges. Anions have negative charges. What force draws the two together?
- 5) Magnetic Attraction WHY?
 - 6) Ionic Attraction
 - 7) Electromagnetic attraction
 - 8) Electrostatic attraction

Exit Ticket:

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