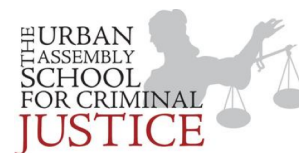


Name: _____ Date: _____

Chemistry ~ Ms. Hart

Class: _____ Anions or Cations



5.12 Intermolecular Forces – Exit Ticket

Part 1: Write whether the molecule is polar or nonpolar. Place an X in the box that shows the type of IMF for this molecule.

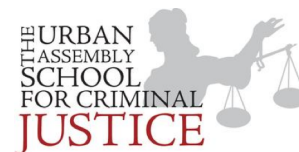
Substance	Polar or Nonpolar	London Dispersion Forces	Dipole-Dipole Forces	Hydrogen Bonding
$\text{H}-\ddot{\text{F}}:$				
$ \begin{array}{ccccc} & \text{H} & & \text{H} & & \text{H} \\ & & & & & \\ \text{H} & -\text{C} & - & \text{C} & - & \text{C} & -\text{H} \\ & & & & & \\ & \text{H} & & \text{H} & & \text{H} \end{array} $				
$ \begin{array}{c} \cdot\cdot \\ \text{H} - \text{N} - \text{H} \\ \\ \text{H} \end{array} $				

Part 2: Explain why HF has a higher boiling point than Cl_2

Name: _____ Date: _____

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