

Name: _____ Date: _____

Chemistry ~ Ms. Hart

Class:

Anions or Cations



2.14 Colligative Properties

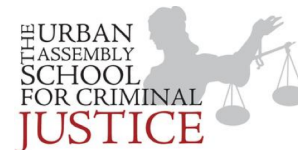
1. Compared to pure water, a solution of calcium chloride has a
 - (1) higher boiling point and higher freezing point
 - (2) higher boiling point and lower freezing point
 - (3) lower boiling point and higher freezing point
 - (4) lower boiling point and lower freezing point
2. When ethylene glycol (an antifreeze) is added to water, the boiling point of the water
 - (1) decreases, and the freezing point decreases
 - (2) decreases, and the freezing point increases
 - (3) increases, and the freezing point decreases
 - (4) increases, and the freezing point increases.
3. Ethylene glycol (antifreeze) is added to cars. Why might that be?
4. At standard pressure, a solution of sugar has a boiling point
 - (1) greater than 100°C and a freezing point greater than 0°C.
 - (2) greater than 100°C and a freezing point less than 0°C.
 - (3) less than 100°C and a freezing point greater than 0°C.
 - (4) less than 100°C and a freezing point less than 0°C.

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1. Compared to pure water, a solution of calcium chloride has a
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 - (2) higher boiling point and lower freezing point
 - (3) lower boiling point and higher freezing point
 - (4) lower boiling point and lower freezing point
2. When ethylene glycol (an antifreeze) is added to water, the boiling point of the water
 - (5) decreases, and the freezing point decreases
 - (6) decreases, and the freezing point increases
 - (7) increases, and the freezing point decreases
 - (8) increases, and the freezing point increases.
3. Ethylene glycol (antifreeze) is added to cars. Why might that be?
4. At standard pressure, a solution of sugar has a boiling point
 - (5) greater than 100°C and a freezing point greater than 0°C.
 - (6) greater than 100°C and a freezing point less than 0°C.
 - (7) less than 100°C and a freezing point greater than 0°C.
 - (8) less than 100°C and a freezing point less than 0°C.