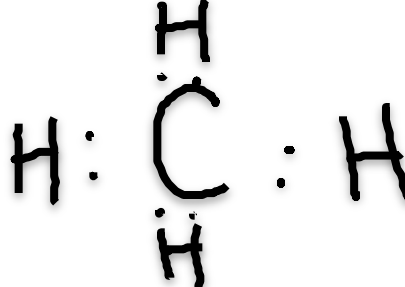


Unit 12

Class Work



NAME

5/21/14

12.1 Hydrocarbons

$$4 + 4 = 8$$

SPARK

→ bonding e^-

Draw the Lewis Dot diagram for CH_4

$VE = 1^{\text{st}}$ # of electron

All elements in group
have same VE Con fig

Objective

SWBAT name and draw hydrocarbons

Organic compounds

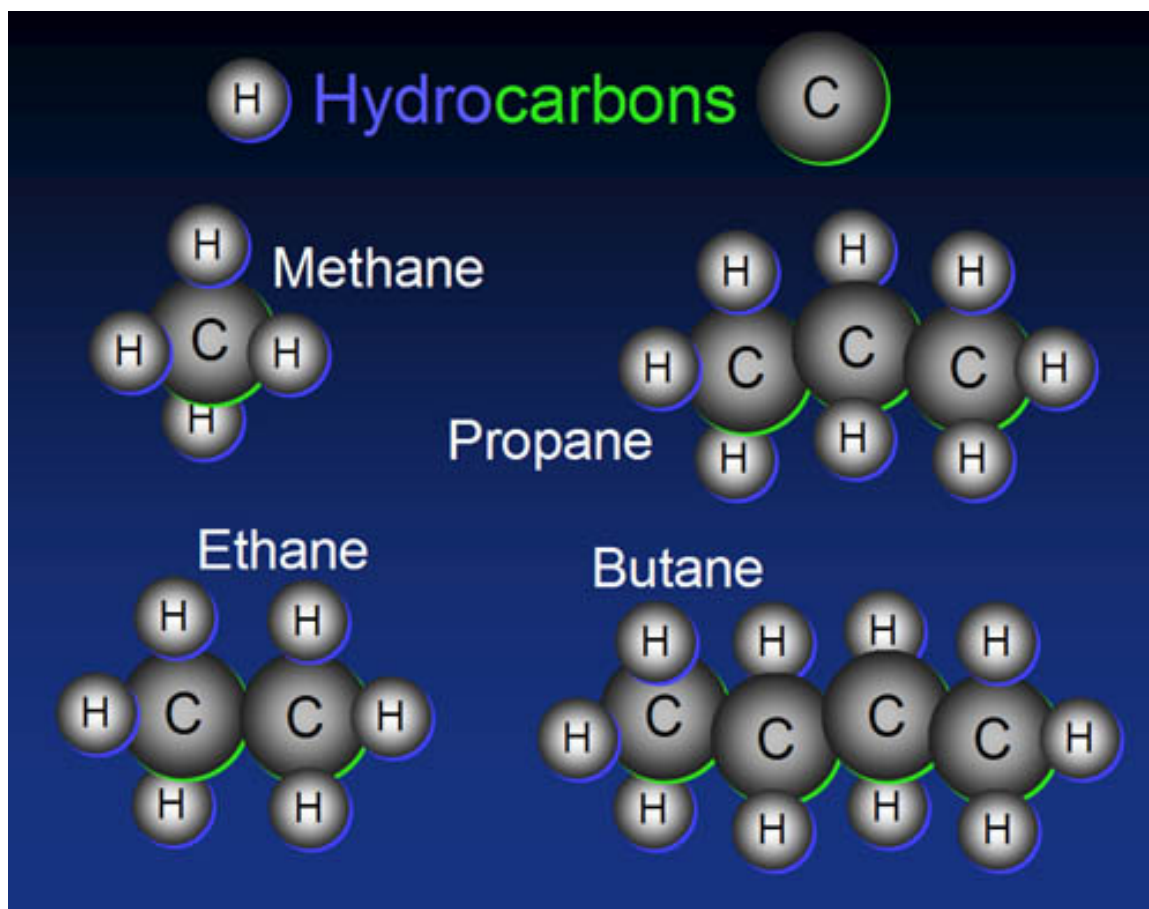
- Contain carbon!
- (~8 million, compared to 100,000 inorganic)
- Each carbon has 4 bonds
- Carbon can make single, double, or triple bonds with itself.



Objective: SWBAT name and draw hydrocarbons

Hydrocarbons

- Compound containing only C and H

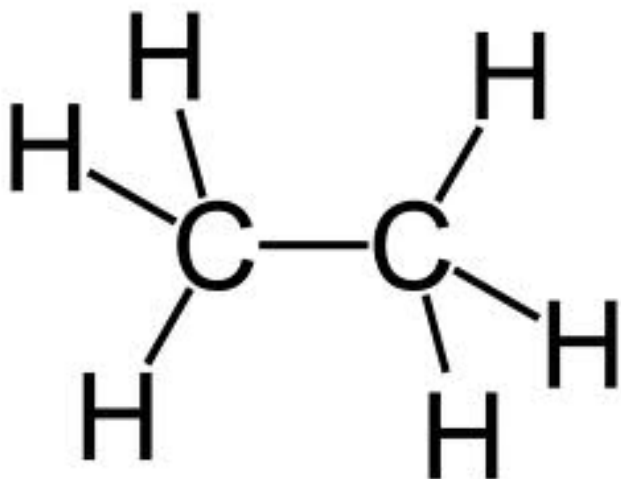


Objective: SWBAT name and draw hydrocarbons

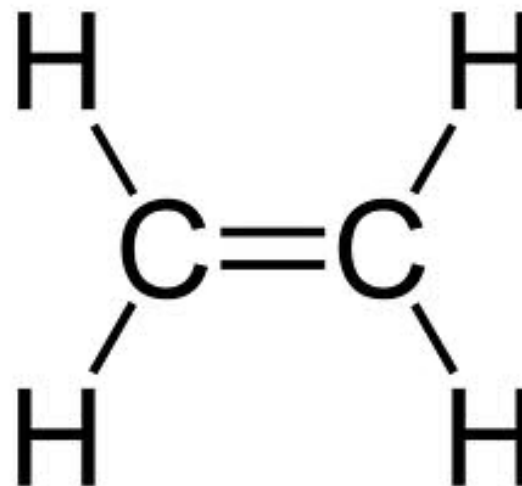
Types of Bonds



- Saturated hydrocarbons contain only single carbon-carbon bonds.



- Unsaturated hydrocarbons contain at least one *multiple* carbon-carbon bond.



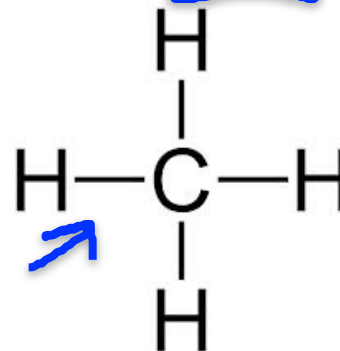
Types of Organic Compounds

Alkanes:

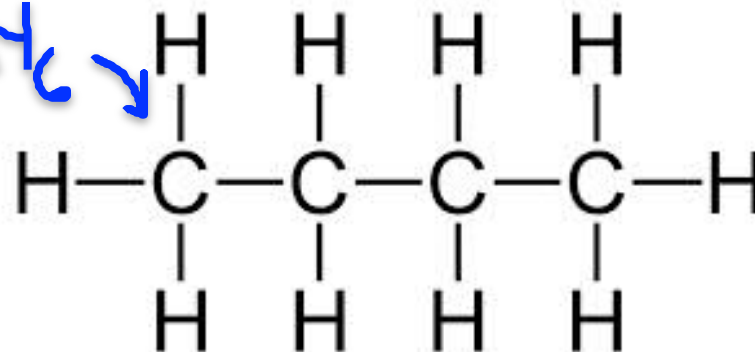
- All single bonds
- Name ends with -ane
- General formula is $C_n H_{2n+2}$
- Differs by CH_2 unit
- Saturated, all single bonds

- Turn to TABLE P

Example: Methane



Butane



Objective: SWBAT name and draw hydrocarbons

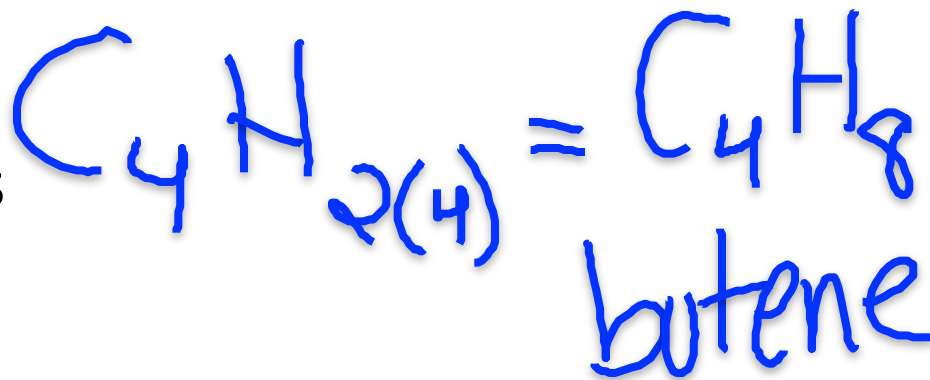
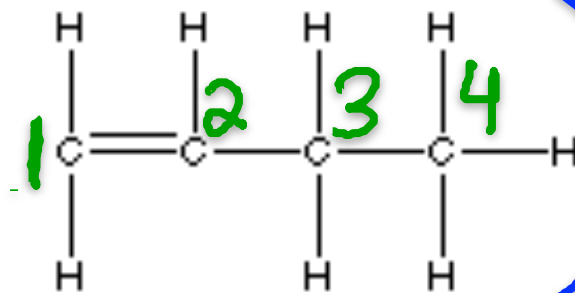
Types of Organic Compounds

1-pentene

Example: 1-butene

Alkenes:

- Contains a double bond
- Name ends with -ene.
- General formula is C_nH_{2n}
- Differs by CH_2 unit
- Unsaturated, contains multiple bond



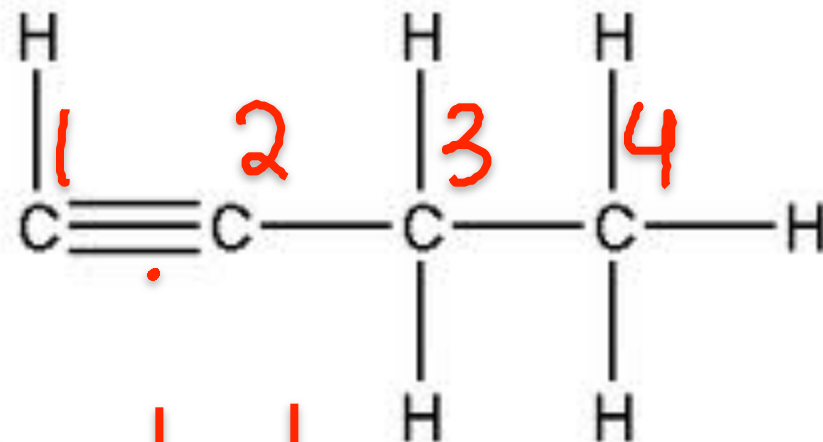
2-hexyne

Types of Organic Compounds

Alkynes:

- Contains a triple bond
- Name ends with -yne
- General formula $C_n H_{2n-2}$
- Differs by a CH_2 unit
- Unsaturated, contains a multiple bond

Example: 1-butyne



1-butyne

Practice

- Draw the straight chain structure for C_4H_{10}

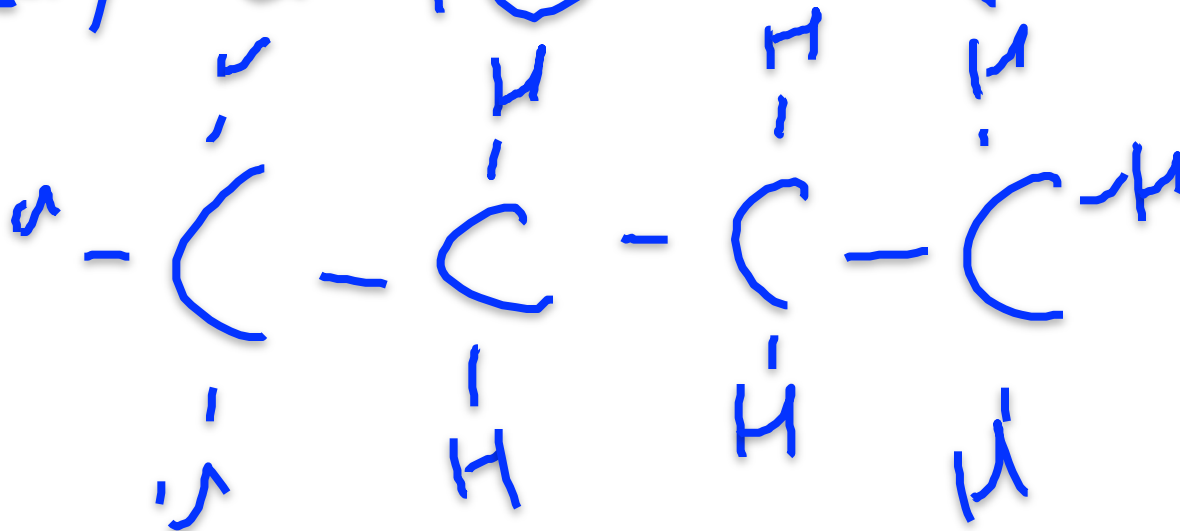
but

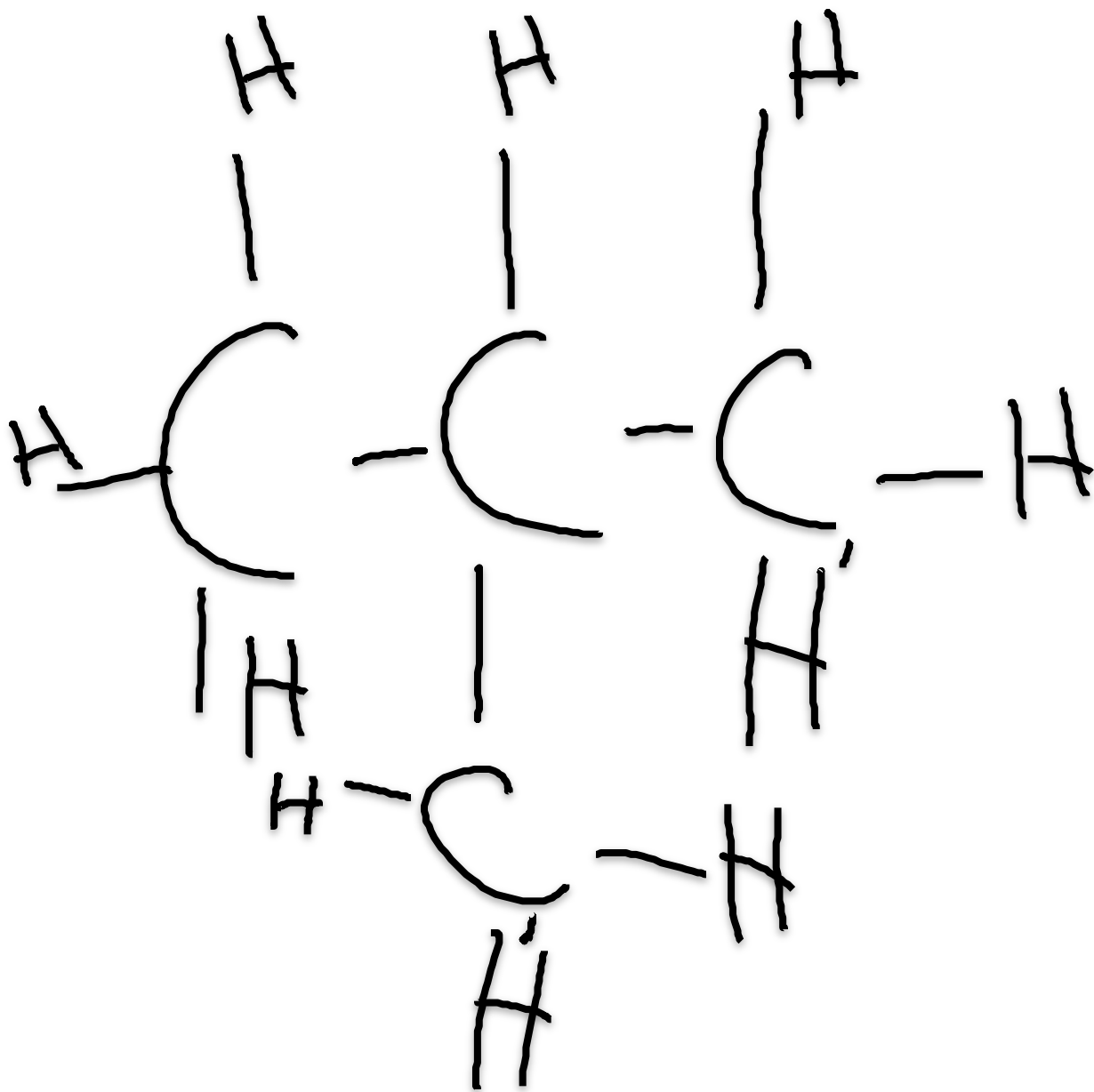
$n=4$

butane

$$4(2) + 2 = 10$$

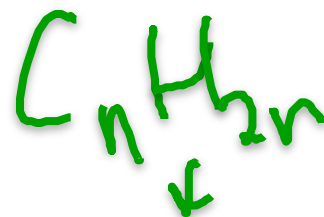
alkane



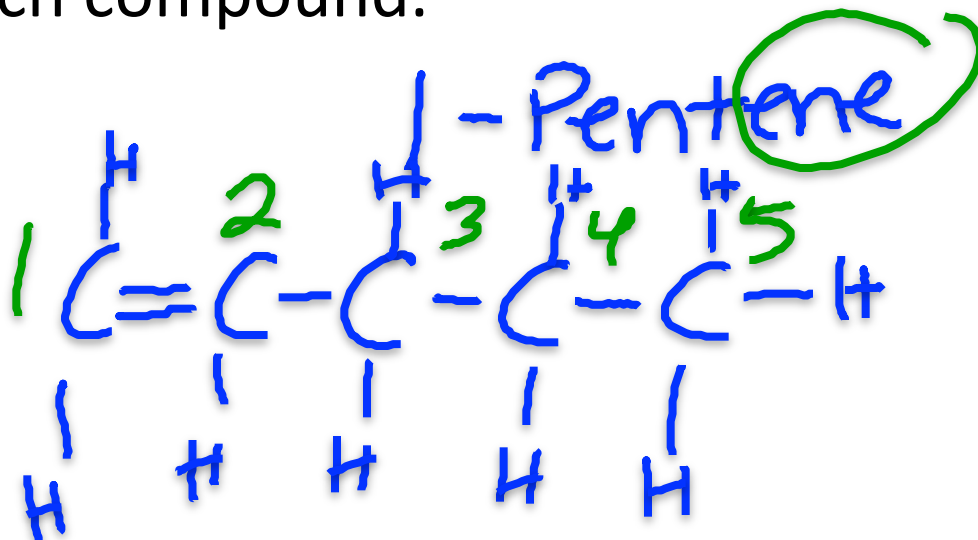


Practice

All of the



- Draw the ~~the~~ straight chain structure for C_5H_{10} , C_6H_{10} , and then name each compound.



Homework

- Complete the rest of the 12.1 HW

Objective: SWBAT name and draw hydrocarbons