Unit 10 Class Work NAME 5/6/14

10.6 Half Life

SPARK

- 1. What is the mass of an alpha particle?
- 2. Write the nuclear decay equation for C-14

Objective

SWBAT complete calculations using the half life

Half Life Lab

- Follow the directions on your lab sheet.
- When you are finished, return all pennies to the ziploc bag and complete graph on the back

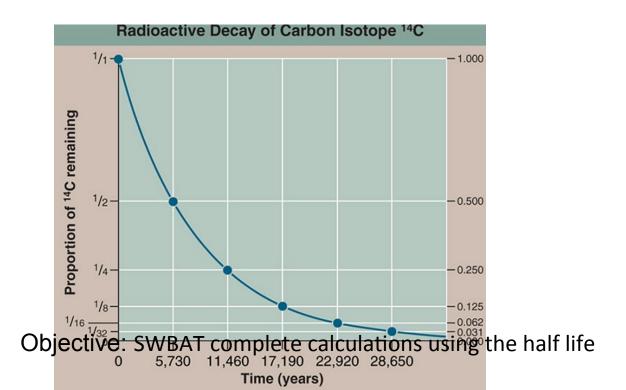
Half Life

Time required for one half of a substance's initial amount to decay

- Examples:
 - Carbon-14 = 5715 years
 - -I-131 = 8.021 days
 - $-U-238 = 4.47 \times 10^9 \text{ years}$

Half Life Graphs

 Graphing half-life shows that the substance decay is not a straight line (as you hopefully saw in your penny graph).



Why do we care?

 Radioisotopes with long half lives are used to date dead organisms.

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Method	Parent/Daughter Isotopes	Half-Lives	Materials Dated	Age Dating Range
Carbon (C)/Nitrogen (N)	C-14/N-14	5,730 yrs.	Shells, limestone, organic materials	100-50,000 yrs.
Potassium (K)/Argon (Ar)	K-40/Ar-40	1.3 billion yrs.	Biotite, whole volcanic rock	100,000-4.5 billion yrs.
tubidium (Rb)/Strontium (Sr)	Rb-87/Sr-87	47 billion yrs.	Micas	10 million-4.5 billion+ yrs.
Uranium (U)/Lead (Pb)	U-238/Pb-206	4.5 billion yrs.	Zircon	10 million-4.5 billion+ yrs.
Uranium (U)/Lead (Pb)	U-235/Pb-207	710 million yrs.	Zircon	10 million-4.5 billion+ yrs.

•Radioisotopes with short half lives are used to fight medical disorders.

Time (days)

Practice!

Using doc camera!

Practice

- Complete the rest of 10.6 Practice Sheet
- Complete the conclusion for Lab #24 (place in classwork bin!)

HOMEWORK

Work on rough draft of Nuclear Power Paper Study ALL notes from Unit 10! Complete online practice quizzes/review vocabulary!